

TERNARY SURFACE COMPLEX: COADSORPTION OF Cu(II), Zn(II), Cd(II) AND NITRILOTRIS(METHYLENE PHOSPHONIC) ACID ONTO BOEHMITE

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This work studies the effect of NTMP (nitrilotris(methylenephosphonic acid)) on the adsorption of Cu(II), Zn(II), and Cd(II) onto boehmite in the pH range 5-9.5. The data are analyzed using the 2-pK constant capacitance model (CCM) assuming ternary surface complex formation. In stoichiometric condition NTMP is more effective to remove Cu(II) than Zn(II) from solution and the contribution of ternary surface complexes are important to model the both metal adsorption. In nonstoichiometric condition and high surface loading with Me(II)/NTMP ratio 1:5 the Cu(II) and Zn(II) adsorption is significantly suppressed. In the case of Cd(II) the free metal adsorption become the more important species.

Keywords: boehmite; Me(II)-phosphonate adsorption; ternary surface complexes.

INTRODUCTION

Metal contaminations still present a major environmental problem.¹ They are introduced in aquatic systems as a result of the weathering of soils and rocks, from volcanic eruptions, and from a variety of human activities.²⁻⁴ The biggest source of heavy metal pollution definitely is industrialization.^{5,6} Furthermore high concentrations of heavy metals enter the aquatic system by release of industrial cooling water into rivers, by dumping sewage sludge and by wet and dry deposition.

Complexation can drastically alter the adsorption characteristics, and hence the mobility, of metals in the environment. The importance of such effects depends on the concentrations of metal-ligand complexes in environmental media and their inherent reactivity.⁷⁻¹¹

In many instances, the aqueous concentration of a metal is controlled by adsorption reactions with the surrounding geochemical phases. The formation of metal-ligand complexes can prevent metal adsorption onto the solid matrix, keeping the metal in the dissolved phase and rendering it highly mobile.^{12,13} Since, metals, ligands, and oxide surfaces can all possess multiple sites of coordination; the possibility exists for the creation of a ternary complex.¹⁴⁻²¹

Therefore it is very important to understand under which conditions heavy metals are mobilized or immobilized. The overall purpose of this research was to study the adsorption of NTMP (nitrilotris(methylenephosphonic acid)), and Me-NTMP using boehmite as a solid phase and to focus especially on the influence of this phosphonate on the adsorption behavior of Cu(II), Zn(II), and Cd(II). To elucidate the competitive effects among boehmite, metals and NTMP; and the possible use of the constant capacity double layer model (CCM) to predict adsorption from double-solute systems using equilibrium constants obtained from single-solute systems was carried out a study of adsorption of these metal ions in presence of NTMP. Were specifically investigated the effects of environmental conditions such as pH variation, the concentration of solutes and various metal-ligand relationships.

EXPERIMENTAL

Materials

The boehmite used in this study was obtained in our laboratory

as was described by Zenobi *et al.*¹⁹ The surface area was determined to be 200 m²/g as measured by BET N₂ adsorption using a Quantachrome NOVA 1200e analyzer. NTMP was obtained from Aldrich Chemicals.

All the solutions were prepared using reagent-grade chemicals and deionized distilled and boiled water. Metal ion solutions were prepared by dissolving the metal nitrate salts in distilled water.

Adsorption experiments

In all cases, the adsorption experiments were carried out in a sealed cylindrical beaker provided with a thermostat water jacket and a total volume of 50 mL.

Me(II) solutions at concentrations of 1 x 10⁻⁵, 5 x 10⁻⁵ and 1 x 10⁻⁴ M and NTMP solutions at concentrations of 5 x 10⁻⁵, 1 x 10⁻⁴ and 5 x 10⁻⁴ M were used for the adsorption study. The Me(II)-NTMP solutions were pre-equilibrated prior to addition to the boehmite suspension. The boehmite concentration used in all experiments was 1 g/L. NaNO₃ was employed to provide an ionic strength of 0.1 M. Carbon dioxide was excluded by purging the suspension with N₂. Suspension pH was adjusted by adding trace amount of HNO₃ or NaOH to keep the pH value in the range 5 to 9.5. NTMP and Me(II) adsorption was calculated from the difference between the total added ligand and metals and that measured in the supernatant after 2 h of equilibrium at 30 ± 0.2 °C. The suspensions were filtered through a 0.45 µm cellulose nitrate filters (Osmonics) prior to analysis.

Dissolved Cu, Zn and Cd concentrations in acidified supernatant were measured by ICP-OES (inductively coupled plasma) emission spectrometry with a Shimadzu Sequential 1000 model III atomic emission spectrometer.

The remaining NTMP in the supernatant was measured as total phosphorous after digestion for 2 h using potassium peroxodisulfate at 100 °C by the molybdenum blue colorimetric test at 880 nm with a UV-V Cecil 2021 spectrometer.

Modeling adsorption data

All experimental data were analyzed using the 2-pK constant capacitance model (CCM). The concepts and assumptions underlying the model have been already discussed in other works.²²⁻²⁴

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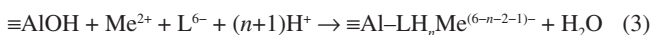
The model describing the acid/base reactions at the boehmite surface is given by the following reactions, intrinsic constants, surface site density and specific capacitance:



$$C = 0.92 \text{ F/m}^2 \quad S_1 = 1.7 \text{ sites/nm}^2$$

These data were taken from Laiti and Öhman.²⁵

Model calculations and optimization of equilibrium constants for adsorption reactions were performed with the nonlinear least-squares optimization program FITEQL 3.2. Technical details of this program were provided by Westall and Herbelin.²⁶ The values of these stability constants were used to adjust the different experimental adsorption edges. The data were modeled considering only mononuclear complex formation for Me(II), NTMP, and Me(II)-NTMP adsorption. The following generalized reactions represent adsorption stoichiometries:



where L^{6-} is the deprotonated ligand (NTMP^{6-}) and Me is the divalent metal. Table 1 gives the surface complexation reactions and the values of the equilibrium constants used for modeling the experimental data. Equilibrium constants for the selected surface reactions were optimized for each data set, and this unique set of constants was then used for all model simulations.

These surface constants were entered into a chemical speciation program MINEQL+ 3.0 to predict NTMP, Me(II), and Me(II)-NTMP sorption.²⁷

Table 1. Equilibrium constant describing the adsorption of Me, NTMP, and Me complexes of NTMP to boehmite, fitted by FITEQL ($I = 0.1 \text{ M NaNO}_3, 30 \text{ }^\circ\text{C}$)

| Surface complexation reactions | log K |
|--|--------------------------|
| $\equiv\text{Al-OH} + \text{L}^{6-} + \text{H}^+ \rightleftharpoons \equiv\text{Al-L}^5 + \text{H}_2\text{O}$ | 24.25 |
| $\equiv\text{Al-OH} + \text{L}^{6-} + 2 \text{H}^+ \rightleftharpoons \equiv\text{Al-LH}^4 + \text{H}_2\text{O}$ | 30.05 |
| $\equiv\text{Al-OH} + \text{L}^{6-} + 3 \text{H}^+ \rightleftharpoons \equiv\text{Al-LH}_2^{3+} + \text{H}_2\text{O}$ | 35.65 |
| $\equiv\text{Al-OH} + \text{L}^{6-} + 4 \text{H}^+ \rightleftharpoons \equiv\text{Al-LH}_3^{2+} + \text{H}_2\text{O}$ | 40.45 |
| $\equiv\text{Al-OH} + \text{Cu}^{2+} \rightleftharpoons \equiv\text{Al-OCu}^+ + \text{H}^+$ | -0.132±0.04 ^a |
| $\equiv\text{Al-OH} + \text{Cu}^{2+} + \text{L}^{6-} + \text{H}^+ \rightleftharpoons \equiv\text{Al-LCu}^{3+} + \text{H}_2\text{O}$ | 30.71 |
| $\equiv\text{Al-OH} + \text{Cu}^{2+} + \text{L}^{6-} + 2 \text{H}^+ \rightleftharpoons \equiv\text{Al-LHCu}^{2+} + \text{H}_2\text{O}$ | 36.71 |
| $\equiv\text{Al-OH} + \text{Zn}^{2+} \rightleftharpoons \equiv\text{Al-OZn}^+ + \text{H}^+$ | -1.541±0.07 ^a |
| $\equiv\text{Al-OH} + \text{Zn}^{2+} + \text{L}^{6-} + \text{H}^+ \rightleftharpoons \equiv\text{Al-LZn}^{3+} + \text{H}_2\text{O}$ | 29.70 |
| $\equiv\text{Al-OH} + \text{Cd}^{2+} \rightleftharpoons \equiv\text{Al-OCd}^+ + \text{H}^+$ | -2.441±0.05 ^a |
| $\equiv\text{Al-OH} + \text{Cd}^{2+} + \text{L}^{6-} + \text{H}^+ \rightleftharpoons \equiv\text{Al-LCd}^{3+} + \text{H}_2\text{O}$ | 27.00 |
| $\equiv\text{Al-OH} + \text{Cd}^{2+} + \text{L}^{6-} + 2 \text{H}^+ \rightleftharpoons \equiv\text{Al-LHCd}^{2+} + \text{H}_2\text{O}$ | 33.00 |
| $\equiv\text{Al-OH} + \text{Cd}^{2+} + \text{L}^{6-} + 3 \text{H}^+ \rightleftharpoons \equiv\text{Al-LH}_2\text{Cd}^+ + \text{H}_2\text{O}$ | 38.20 |

^aTaken from ref. 19

Stability constants for NTMP protonation and for Me(II)-NTMP complex formation were taken from Deluchat *et al.* (Table 2).²⁸ NTMP, Me(II), and Me(II)-NTMP adsorption data were modeled by defining the minimum number of surface species that consistently yielded a good fit to all data.

Table 2. Speciation of NTMP and metal complexes of NTMP^a

| Protonation constants (0.1 M NaNO ₃ , 30 °C) ^a | | | |
|---|--------------------|-------|-------|
| log K ₁ | | 12.50 | |
| log K ₂ | | 7.22 | |
| log K ₃ | | 5.90 | |
| log K ₄ | | 4.59 | |
| log K ₅ | | 1.6 | |
| log K ₆ | | 0.5 | |
| Complexation constants (0.1 M NaNO ₃ , 30 °C) ^b | | | |
| | | log K | log β |
| Cu(II) | MeL | 17.40 | |
| | MeHL | 6.35 | 23.78 |
| | MeH ₂ L | 4.57 | 28.36 |
| | MeH ₃ L | 3.46 | 31.8 |
| | MeH ₄ L | 1.4 | 33.3 |
| Zn(II) | MeL | 16.3 | |
| | MeHL | 6.1 | 22.5 |
| | MeH ₂ L | 4.89 | 27.3 |
| | MeH ₃ L | 4.0 | 31.38 |
| | MeH ₄ L | 2.5 | 33.8 |
| Cd(II) | MeL | 12.2 | |
| | MeHL | 7.16 | 19.4 |
| | MeH ₂ L | 5.68 | 25.0 |
| | MeH ₃ L | 4.13 | 29.2 |

^a i is the number of acid functions. $K_i = [\text{H}_i\text{L}]/\{[\text{H}^+][\text{H}_{i-1}\text{L}]\}$; ^b $K = [\text{MeH}_i\text{L}]/\{[\text{MeH}_{i-1}\text{L}][\text{H}^+]\}$; $\beta = [\text{MeH}_i\text{L}]/\{[\text{Me}][\text{L}][\text{H}^+]\}$. Taken from ref. 28

RESULTS AND DISCUSSION

In these experiments both concentrations NTMP and Me(II) were measured. The experimental data of the amounts of phosphonic acid and metal adsorbed obtained for every Me(II)-NTMP-combination were plotted separately so, we can see the influence of the phosphonic acid on the metal adsorption and vice versa.

The adsorption studies of Me (II)-NTMP were performed in stoichiometric and non-stoichiometric conditions.

Stoichiometric conditions

The adsorption edges for both system Cu(II)-NTMP and Zn(II)-NTMP at $5 \times 10^{-4} \text{ M}$ high solution concentrations are shown in Figure 1. Under these conditions, only Cu(II) and Zn(II) metals are considered. Cd(II) was excluded due to the low solubility of Cd-NTMP complex.

At low pH values a significant increase in Cu(II) and Zn(II) adsorption is noted in presence of NTMP $5 \times 10^{-4} \text{ M}$ (full lines in Figure 1) as compared with Me(II) adsorbed onto boehmite in absence of NTMP (open circles Figure 1). The ligand effect is more pronounced for Zn(II) than for Cu(II), this is related to the values of the constants of both complexes in solution (Table 2). At $\text{pH} \approx 6, 7$ the adsorption of both Cu(II) and Zn(II) reach a maximum and then decrease as pH increase. At low pH values the model predicts a higher affinity of the boehmite surface by complexed Me(II) rather than the metal ion in its free form. Also the model prediction shows that at higher pH adsorption of free metal ($\equiv\text{AlOMe}^+$) is the dominant species. However, we can see that from pH 7 the adsorbed Cu exists mainly in its free form while, in the case of Zn at pH about 7.5 an equal contribution from both free and complexed forms is observed.

For the present study two mononuclear ternary surface complexes for Cu-NTMP system ($\equiv\text{AILHCu}$ and $\equiv\text{AILCu}$) and only one for Zn-

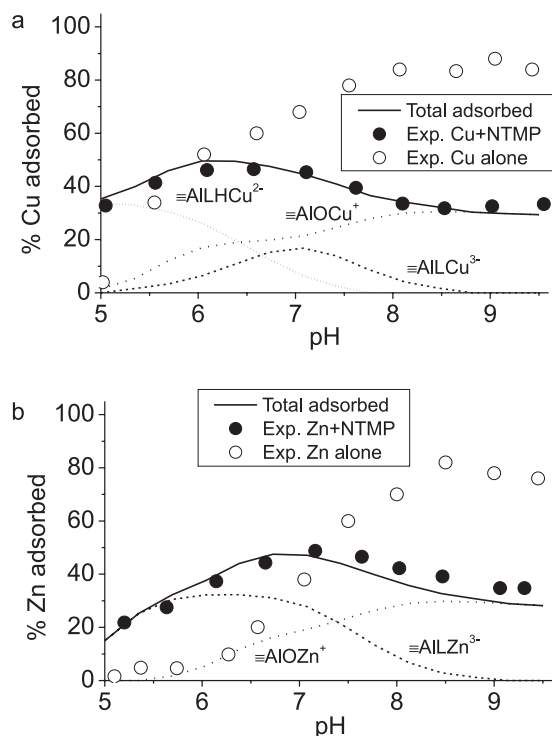


Figure 1. Percentage Cu(II) and Zn(II) 5×10^{-4} M adsorbed vs pH in the presence of NTMP 5×10^{-4} M. The curves are calculated using the constants from Table 1. Conditions: 1 g/L boehmite, 0.1 M NaNO₃

NTMP system ($\equiv\text{AILZn}^+$), in addition to adsorption of both Me(II) and NTMP in its free form had to be considered to obtain a good fit of the experimental data. Both experimental data and model results suggest that at low pH the ternary surface complexes are predominant whereas binary complexes Me(II)-AIOH and NTMP-AIOH largely predominate at high pH.

Both Cu (II) and Zn (II) metals in the presence of 1×10^{-4} M NTMP (Figure 2Sb, supplementary material) show a typical adsorption edge of bell-shaped. In both cases the maximum in the metal adsorbed is close to 100%. It can be seen that 50% of the total available Cu(II) is immobilized by the presence of NTMP at pH 5 but, only 20% of total Zn(II) concentration is attached to surface at this pH. At lower pH the effect of ternary surface complexes in the adsorption process is higher for Cu than for Zn and the model predicts two ternary surface complexes for the Cu(II) adsorption and only one is necessary to fit the adsorption results in the case of Zn(II). In both cases after reaching a maximum around pH 7 the metal binary species ($\equiv\text{AIOMe}^+$) dominates the adsorption.

The model predicts that Cd is adsorbed in its free form preferment and the decrease observed from pH 8 is lower than for Cu(II) and Zn(II).

The adsorption decrease exhibited by the three metals at the higher pH is due to a competitive effect between the active surface sites and the NTMP in solution for the metal ion. The immobilization of these metal ions at high pH is mainly regulated by the Me(II)-NTMP constants values (Table 2).

When Me(II)-NTMP concentration is 5×10^{-5} M, Me(II) and NTMP are adsorbed mainly in their free forms (Figures 3Sa and 3Sb, supplementary material). In the presence of an excess of surface sites at pH 5 only 20% of total Cu is removed while, at the same pH the NTMP is completely adsorbed. The adsorption edge of Zn(II) in 5×10^{-5} M Zn(II)-NTMP system exhibits the same trend as the free Zn(II) adsorption pattern. However, in both systems the contribution

of ternary surface complexes must be invoked to adjust the experimental results of Cu(II), Zn(II) and NTMP adsorption. Due to the high stability of the complexes Cu-NTMP and Zn-NTMP in solution, at pH greater than 8 only a slight decrease in the adsorption of Cu(II) and Zn(II) is caused by the presence of NTMP. When concentration of the solution Cd-NTMP is 5×10^{-5} M the model predicts that both Cd and NTMP adsorbs only in its free form.

Non-stoichiometric conditions

Deficiency of surface sites

Experimental and model results for Me(II) 1×10^{-4} and NTMP 5×10^{-4} M adsorption onto boehmite are shown in Figure 2. Because metal concentration is five times lower than the concentration of NTMP is expected that the metal is completely complexed with NTMP in solution while the NTMP will be mostly in its free form. For the system Cu-NTMP experimental results indicate a poor adsorption of the metal in the entire pH range reaching a maximum of 20% at the higher pH. In the case of Zn-NTMP the metal adsorption is even less not exceeding 10%. With the same surface complexation constants used to fit the data in stoichiometric conditions the

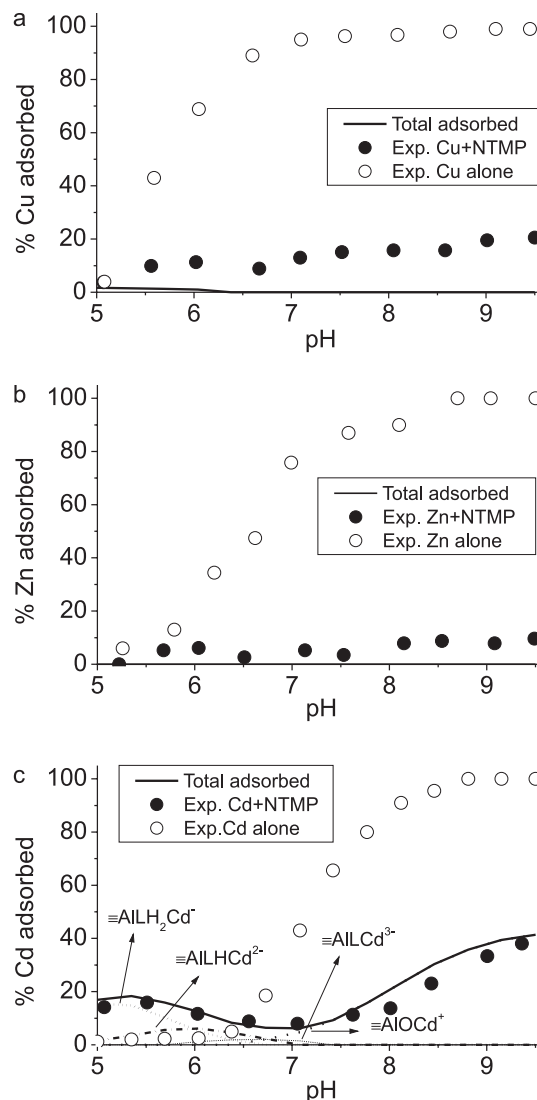


Figure 2. Percentage Cu(II), Zn(II) and Cd(II) 1×10^{-4} M adsorbed vs pH in the presence of NTMP 5×10^{-4} M. The curves are calculated using the constants from Table 1. Conditions: 1 g/L boehmite, 0.1 M NaNO₃

model predicts that the percentage of Cu (II) and Zn (II) adsorbed is practically negligible.

Deluchat *et al.* and Sawada *et al.* calculated the complexation constants of different NTMP-Me complexes with a ratio ligand / metal (1/1), no data in the literature that addresses other molar relationships.^{28,29} It is possible that the relationship used in this study, with a ligand concentration of five times greater than the concentration of metal contribute to form other complexes such as M_2L or ML_2 not taken into account when modeling the experimental data leading to the observed deviations between the experimental results and those arising from the model, especially in the case of NTMP-Cu.

For Cd 1×10^{-4} and NTMP 5×10^{-4} M the model predicts very satisfactorily the adsorption of both solutes NTMP and Cd (II) over the entire pH range (Figure 4Sa to compare the NTMP behavior when this ligand is alone). Besides we can see, unlike the Cu (II) and Zn (II) that surface ternary species contribute to the adsorption of both NTMP and Cd (II). Another important aspect is that besides the species $\equiv AlLCd^{3-}$ and $\equiv AlLHCd^{2-}$ already considered in the stoichiometric conditions was necessary to introduce the species $\equiv AlLH_2Cd$ to achieve a good fit to the data. We would assume that in the case of Cd (II) no change in the species in solution for the ligand excess, so that all experiences in stoichiometric and non stoichiometric conditions could be adjusted by using the same set of constants.

Excesses of surface sites

In this case where the L/M is also 5/1, but at concentrations ten times lower than those considered in the previous paragraph, we see that the model adequately predicts the adsorption behavior of both ligand and metal (Figures 3 and 5Sa, supplementary material).

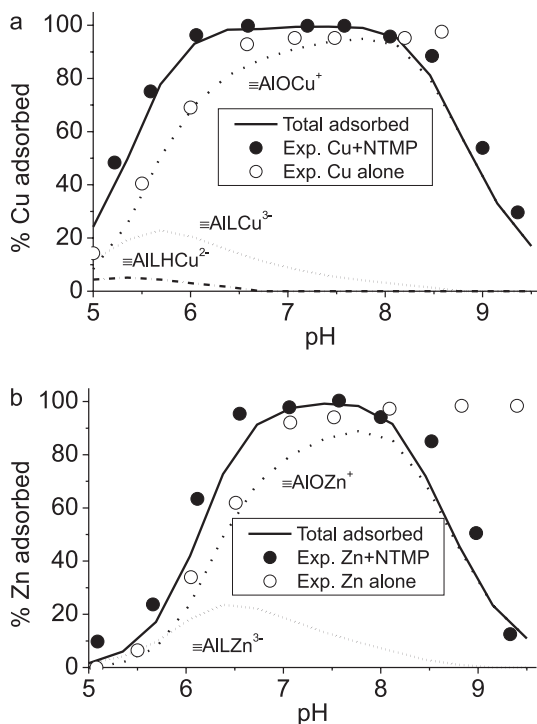


Figure 3. Percentage Cu(II) and Zn(II) 1×10^{-5} M adsorbed vs pH in the presence of NTMP 5×10^{-5} M. The curves are calculated using the constants from Table 1. Conditions: 1 g/L boehmite, 0.1 M $NaNO_3$

At low surface covering an excess of NTMP has not the same impact as in the case of a deficiency of surface sites. A slight increase in adsorption is observed for the case of Cu (II) at pH below 6.5 where

the ternary surface species presents an important contribution to the total metal adsorption.

Below pH 8 the Zn (II) and Cd (II) (Figure 5Sb, supplementary material) did not significantly alter their behavior by the presence of NTMP giving a similar adsorption pattern with a slight increase in metal retention.

At pH 8 begins to decrease markedly the adsorption of Cu (II) and Zn (II) by effect of ligand excess, leading to the same percentages of adsorption found at more acidic pH. As in previous cases this is attributed to the increased stability of the complexes Me-NTMP in alkaline media and the lower affinity of Me-NTMP complexes by the oxide surface in this condition.

On the other hand, Cd (II) (Figure 5Sb, supplementary material) not requires the presence of ternary surface complexes to fit the experimental data showing the same profile as the free metal adsorption. Again here the adsorption behavior is the result of the consequent competition between chelation and adsorption.

In both stoichiometric and non-stoichiometric conditions, the presence of metals has no significant effect on the NTMP adsorption edge. The adsorption level reaches the pattern followed by the free NTMP. However, it is necessary to consider the formation of ternary surface complexes to explain the phosphonate adsorption in the presence of metal ions (Figures 1Sa to 5Sa, supplementary material).

CONCLUSIONS

This study provides significant information about adsorption from Me(II)-NTMP mixtures in the pH-range of natural waters (5-9.5) and at different adsorbent concentrations. The 2-pK constant capacitance model (CCM) agreed with the overall adsorptive trends for Cu(II), Zn(II) and Cd(II) in presence of stoichiometric and nonstoichiometric concentrations of NTMP at different surface coverage conditions. Experimental and model results suggest the adsorption of free metal ion and NTMP as well as ternary complexes $\equiv Al-NTMP-Me(II)$. These ternary surface complexes produce drastic changes in the adsorption trend for Cu(II), Zn(II) and Cd(II). These results are in agreement with the statement by Zenobi and Rueda, who demonstrated that the presence of 1-hydroxyethane-(1,1-diphosphonic acid) (HEDP) in natural waters could aid in the removal of those divalent metals dissolved.²⁰ However there is a substantial difference in the complexing behavior between HEDP and NTMP in conditions of high phosphonate/metal ratio and high surface loading. NTMP is a stronger complexing agent than HEDP and this fact may have an important environmental impact since it might contribute to mobilize Cu(II), Zn(II) and Cd(II) from sediments.

SUPPLEMENTARY MATERIAL

The Figures 1Sa to 5Sa (Percentage NTMP adsorbed vs pH in the presence of Cu(II), Zn(II) and Cd(II)), Figures 2Sb and 3Sb (Percentage Cu(II), Zn(II) and Cd(II) adsorbed vs pH in the presence of NTMP) and Figure 5Sb (Percentage Cd(II) 1×10^{-5} M adsorbed vs pH in the presence of NTMP 5×10^{-5} M) are available free of charge at <http://quimicanova.sbq.org.br>, as PDF file.

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